Radiative de-excitation of vibrationally excited C₂⁻

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Hot molecules isolated in the vacuum environment dissipate their internal energy solely though emission of IR photons, which thereby govern the stability and reactivity of such molecules, for example those in the interstellar space. In general, the radiative cooling of hot homonuclear diatomic molecules is very slow because vibrational transitions are dipole forbidden. In fact, hydrogen molecules require approximately 10⁶ s for such transitions¹. C_2^- is an exception that can be cooled much more efficiently via electronic transitions between the ground state $X^2\Sigma_g^+$ and the low-lying excited state $A^2\Pi_u$, as schematically shown in Fig. 1. The timescales of the transitions from X to A and from A to X are theoretically predicted to be on the orders of ms and µs, respectively², whereas these values have not been experimentally confirmed.

vibrational states aided by the PGOPHER code³.

populations at the high vibrational states and increases at the

low vibrational states⁴. The time profiles agree well with the

simulated time evolution of the populations based on the

theoretical rates² and the Boltzmann distribution at the

vibrational temperature of 4500 K in the ion source.

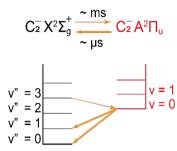


Fig. 1 De-excitation pathways of C_2^- from v'' = 3in the X state.

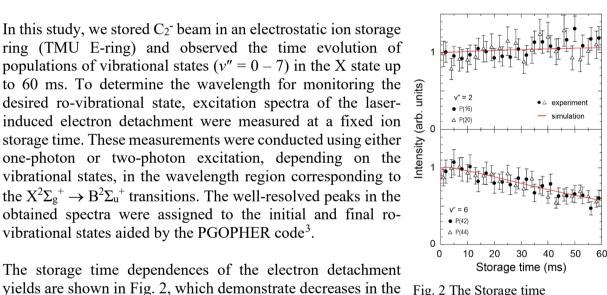


Fig. 2 The Storage time dependences of the electron detachment yields at v'' = 2 (top) and v'' = 6 (bottom) in the X state. Red lines are the simulated time profiles.

¹ Pachucki, K.; Komasa, J. Magnetic dipole transitions in the hydrogen molecule. *Phys. Rev. A* 2011, 83, 032501. ² Rosmus, P.; Werner, H. J. Multireference-CI calculations of radiative transition probabilities in C₂⁻. J. Chem. Phys. 1984, 80, 5085-5088.

³ Western, C. M. A program for simulating rotational, vibrational and electronic spectra. J. Quant. Spectrosc. Radiat. Transfer 2017, 186, 221–242.

⁴ Iida, S. et al., State-Selective Observation of Radiative Cooling of Vibrationally Excited C₂⁻. J. Phys. Chem. Lett. 2020, 11, 10526–10531.